

WILLIAM LE NOBLE.

Jerome A. Berson and Elli S. Hand: Thermal Rearrangements of the 7-Carbomethoxy- Δ^2 -norcarenes.

Page 1982. In Table II, 0.024 in the right-hand column refers to methyl cycloheptanecarboxylate; 1.00 refers to a missing entry, "Others."--JEROME A. BERSON.

Thomas J. Wallace: Reactions of Thiols with Sulfoxides. I. Scope of the Reaction and Synthetic Applications.

Page 2018. In ref. 3, 1044 should read 888. In ref. 6, 504 should read 501.--THOMAS J. WALLACE.

Richard B. Simpson: Association Constants of Methylmercuric and Mercuric Ions with Nucleosides.

Page 2059. The last sentence in the Abstract, "The affinity of calf thymus DNA for inorganic mercury is at least an order of magnitude greater than that of the nucleosides," is not true. H. T. Miles first expressed doubt that a more stable structure than the native nucleic acid should be formed on adding mercury. S. Katz then pointed out that my argument (in the second paragraph of the section on Polynucleotides) involves the implicit assumption that each inorganic mercury reacts with only one base of the nucleic acid, an assumption clearly inconsistent with the release of more than one proton per mercury.1

In the range of 0 to 0.5 mercury added per nucleotide of DNA, it is probable that nearly every mercury reacts with two bases on the same molecule, i.e., double-stranded helix. Then from the following equation it may be seen that the partition of mercury between a couple of bases on the polynucleotide and the same two bases on nucleosides free in solution would be dependent on concentration

 HgB_1B_2 + polynucleotide base pair

polynucleotide base pair: $Hg + B_1 + B_2$

where B_1 and B_2 need not be a Watson-Crick base pair. I.e., The term "relative affinity" has no meaning since the two associations are not strictly analogous. A more familiar example of such an equilibrium occurs if one attempts to compare the association constant of a metal chelate and the association constant of the metal with molecules having just one of the functional groups.

(1) T. Yamane and N. Davidson, J. Am. Chem. Soc., 83, 2599 (1961). RICHARD B. SIMPSON.

CH₂CR₁ Ö TIT

Robert A. Benkeser, Yoichiro Nagai, James L. Noe, Robert F. Cunico, and Peter H. Gund: The Synthesis of Cyclic Silicon Systems.

Page 2218. Formula III should appear as

Pages 2448 and 2450. The name given to structure VI should read 1,1-dichloro-4-bromosila-2-cyclopentene instead of 1,1dichloro-3-bromosila-2-cyclopentene. The structural formula for the compound is correct as given.---R. A. BENKESER.

E. Wenkert, P. W. Jeffs, and J. R. Mahajan: A Hibaene Model.

Robert E. Rinehart and Jack S. Las'xy: The Isomerization of 1,3-Cyclooctadiene to 1,5-Cyclooctadiene via the Rhodium(I) π -Complex.

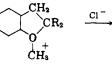
Page 2517. In footnote 12, the infrared band should read 2260 cm. -1.--ROBERT E. RINEHART.

Arthur R. Lepley: *m*-Complex Interactions. II. Methyl-Substituted Benzenes and Naphthalenes with Tetracyanoethylene

Page 2547. Column 1, line 25, should read 2 = 3 = 5 = 6, instead of 2 = 3 = 4 = 5.--ARTHUR R. LEPLEY.

Stephen E. Cantor and D. Stanley Tarbell: The Formation of cis- and trans-Perhydrobenzofurans from 2-(2-Methoxycyclohexyl)ethanol Derivatives. Reactions Proceeding through Methoxyl Participation.

Page 2905. In column 2, structure BB should read



D. S. TARBELL.

P. Veeravagu, R. T. Arnold, and G. W. Eigenmann: Competitive Elimination-Substitution Reactions. Some Dramatic Differences between Bromides and Tosylates.

Page 3072. The third author's correct name is G. W. Eigenmann.

Page 3073. Table II is correctly represented below.

PRODUCT ANALYSIS OF THE REACTION OF ALKOXIDES WITH BROMIDES AND TOSYLATES -Compound RY Elimination, Substitution, Total yield. R ⇒ Y = Alkoxide (1.0 N)% % % NaOCH₃ Ι n-C18H37 Br 1 96 97 (a) $n - C_{18}H_{37}$ Br (b) KO-t-C₄H₉ 85 1297 n-C₁₈H₃₇ NaOCH₃ II OTS 9596 (a) 1 n-C18H37 OTs KO-t-C₄H₉ 99 100 (b) 1 III -CH₂CH₂ Br KO-t-C4H9 73 19 92 IV CH₂CH₂ OTs NaOCH, 84 85 (a) 1 CH₂CH₂ OTs KO-t-C₄H₉ (b) 1 80 81 V CH₂CH₂CH₃ OTs KO-t-C₄H, 1 88 89 VI KO-t-C4H9 100 00 1009 CH₂CH₂ Br VII KO-t-C₄H₉ 1009 CH₂CH₂ OTs 100 00 VIII OTs NaOCH₂ 84 (a)5 79 CH₂CH OTs (b) KO-t-C₄H₉ 7171

TABLE II

E. WENKERT.